

Tris[tris(ethane-1,2-diamine)cobalt(II)] bis[octacyanidomolybdate(V)] dihydrate

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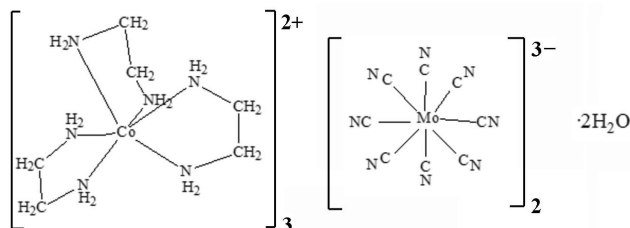
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Key indicators: single-crystal X-ray study; $T = 153$ K; mean $\sigma(\text{C}-\text{C}) = 0.005$ Å; disorder in main residue; R factor = 0.048; wR factor = 0.102; data-to-parameter ratio = 15.1.

In the title compound, $[\text{Co}^{\text{II}}(\text{C}_2\text{H}_8\text{N}_2)_3][\text{Mo}^{\text{V}}(\text{CN})_8]_2 \cdot 2\text{H}_2\text{O}$, $\text{N}-\text{H} \cdots \text{N}$ and $\text{N}-\text{H} \cdots \text{O}$ hydrogen-bonding interactions give rise to a three-dimensional network. In the crystal structure, each Mo polyhedron has a square-antiprismatic shape, while the Co complexes show distorted octahedral geometry with an occupancy of 50%. One of the Co atoms resides on a crystallographic inversion centre.

Related literature

For information on octacyanidometalate-based compounds, see: Bok *et al.* (1975); Lim *et al.* (2006) and literature cited therein; Przychodzeń *et al.* (2006) and literature cited therein; Sieklucka *et al.* (2002); Willemin *et al.* (2003); Withers *et al.* (2006). For related literature, see: Aschwanden *et al.* (1993); Müller *et al.* (2006).



Experimental

Crystal data

 $[\text{Co}(\text{C}_2\text{H}_8\text{N}_2)_3][\text{Mo}(\text{CN})_8]_2 \cdot 2\text{H}_2\text{O}$
 $M_r = 1361.96$

 Monoclinic, $P2_1/n$
 $a = 9.2113$ (3) Å

 $b = 30.5439$ (8) Å

 $c = 11.4022$ (3) Å

 $\beta = 94.138$ (1)°

 $V = 3199.63$ (16) Å³
 $Z = 2$

 Mo $K\alpha$ radiation

 $\mu = 1.20$ mm⁻¹
 $T = 153$ (2) K

 $0.25 \times 0.23 \times 0.18$ mm

Data collection

Rigaku R-Axis Spider diffractometer

Absorption correction: multi-scan (ABSCOR; Higashi, 1995)

 $T_{\text{min}} = 0.753$, $T_{\text{max}} = 0.813$

27139 measured reflections

6276 independent reflections

 4952 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.054$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.047$
 $wR(F^2) = 0.101$
 $S = 1.06$

6276 reflections

415 parameters

56 restraints

H-atom parameters constrained

 $\Delta\rho_{\text{max}} = 0.42$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.55$ e Å⁻³
Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H} \cdots A$	$D-\text{H}$	$\text{H} \cdots A$	$D \cdots A$	$D-\text{H} \cdots A$
N9—H9C \cdots N7 ⁱ	0.92	2.24	3.064 (4)	148
N11—H11C \cdots N7 ⁱ	0.92	2.09	2.985 (4)	165
N12—H12C \cdots N6 ⁱⁱ	0.92	2.10	2.955 (4)	154
N13—H13C \cdots N8 ⁱⁱ	0.92	2.55	3.138 (4)	122
N9—H9D \cdots N2	0.92	2.32	3.117 (4)	145
N14—H14D \cdots N1	0.92	2.28	3.130 (5)	154
N10—H10C \cdots O3 ⁱⁱⁱ	0.92	1.96	2.856 (10)	164
N10—H10D \cdots N5 ⁱⁱⁱ	0.92	2.49	3.125 (5)	126
N11—H11D \cdots N6 ⁱⁱⁱ	0.92	2.58	3.163 (4)	122
N13—H13D \cdots N5 ⁱⁱⁱ	0.92	2.11	3.012 (4)	167
N16—H16C \cdots N4 ^{iv}	0.92	2.52	3.19 (2)	130
N17—H17A \cdots N2 ^v	0.97	2.49	3.29 (2)	140
N20—H20C \cdots N2 ^{iv}	0.85	2.44	3.02 (2)	127
O2—H2A \cdots N8 ^{vi}	0.85 (14)	2.41 (14)	3.207 (9)	156 (12)
O2—H2B \cdots N2 ^{vi}	0.85 (15)	2.50 (15)	3.245 (9)	147 (12)

Symmetry codes: (i) $x - \frac{1}{2}, -y + \frac{1}{2}, z - \frac{1}{2}$; (ii) $x + \frac{1}{2}, -y + \frac{1}{2}, z - \frac{1}{2}$; (iii) $x, y, z - 1$; (iv) $-x + 1, -y + 1, -z + 1$; (v) $x + 1, y, z$; (vi) $-x, -y + 1, -z + 1$.

Data collection: *RAPID-AUTO* (Rigaku, 2004); cell refinement: *RAPID-AUTO*; data reduction: *RAPID-AUTO*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXL97* and *PLATON* (Spek, 2003).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SI2095).

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supplementary materials

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Tris[tris(ethane-1,2-diamine)cobalt(II)] bis[octacyanidomolybdate(V)] dihydrate

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Comment

Recently, octacyanometallates $[M(\text{CN})_8]^{3-/4-}$ ($M = \text{Mo}$ and W) appear as new versatile building blocks and have been investigated extensively (Przychodzeń *et al.*, 2006; Sieklucka *et al.*, 2002). These species might show various geometrical structures (*e.g.*, square antiprism, dodecahedron, bicapped trigonal prism) depending on the external environments. However, only a few examples of cobalt-octacyanometalate coordination networks have been reported until now (Willemin *et al.*, 2003; Przychodzeń *et al.*, 2006 and literature cited therein). In the title bimetallic compound, $[\text{Co}^{\text{II}}(\text{en})_3][\text{Mo}^{\text{V}}(\text{CN})_8]_2 \cdot 2\text{H}_2\text{O}$, (Fig. 1), the Mo^{V} atom is coordinated by eight CN groups with Mo—C distances ranging from 2.128 (4) to 2.179 (4) Å in a distorted square antiprism. The Mo^{V} —C bond distances are comparable to previously reported values (Withers *et al.*, 2006; Lim *et al.*, 2006 and literature cited therein). The geometry around atom Co1 is a distorted octahedron with (D_3 symmetry) with an average Co1—N bond distance 1.972 Å and the N—Co1—N bond angles ranging from 84.34 (12)° to 176.73 (12)°. Among the two independent Co(en)₃ complexes the cobalt atom Co2 resides on a crystallographic inversion centre, resulting in inversion related 50:50 disorder of the chiral Λ and Δ forms. Disorder refinement (Müller *et al.* 2006) was necessary to present a suitable chiral model of the complex. (Fig. 1).

Several classic intermolecular N—H...N hydrogen bonds (Fig. 2) between the non-disordered Co(en)₃ and the Mo(CN)₈ complexes form a complicated three-dimensional network in the structure. The disordered Co2 unit and the disordered water molecules are not considered for further (possible) hydrogen bonding contacts. An interesting example structure with two crystallographically independent Co(en)₃ complexes, different vanadates and six water molecules shows an impressive number of 23 N—H...O and 13 O—H...O hydrogen bonds in the chiral space group P1 (Aschwarden *et al.* 1993). It may be thinkable that the title structure belongs also to a non-centrosymmetric space group (P 21), but in that case with a dominating part of centrosymmetry in data.

Experimental

For the preparation of the title compound, all of the following procedures were carried out in the dark to avoid decomposition of $(\text{Bu}_3\text{NH})_3[\text{Mo}(\text{CN})_8] \cdot 4\text{H}_2\text{O}$ (Bok *et al.*, 1975). Yellow block crystals suitable for X-ray single-crystal structure determination were grown at room temperature by slow diffusion of an aqueous solution of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ (0.3 mmol) and ethane-1,2-diamine (en, 0.9 mmol) and an aqueous solution of $(\text{Bu}_3\text{NH})_3[\text{Mo}(\text{CN})_8] \cdot 4\text{H}_2\text{O}$ (0.2 mmol) for four weeks. The resulting crystals were collected, washed with H_2O and dried in air.

Refinement

All non-H atoms were refined anisotropically. The H atoms on nitrogen atoms were located from the difference Fourier maps, and the H atoms of water molecules were placed in calculated positions. The H atoms of the Co complexes were placed in calculated positions with C—H and N—H distances 0.99 Å and 0.92 Å, respectively, with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C}, \text{N})$.

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The calculation of the H atoms for the disordered Co₂ complex was possible with many FREE instructions using *SHELXL97* (Sheldrick, 2008). For atoms N17 and N20 the four H-atom coordinates from the difference map were fixed with AFIX 3 instructions. These N—H distances range between 0.767 and 0.968 Å. The model refinement of the Co₂(en)₃ complex was controlled with the programme *PLATON* (Spek, 2003), (LATT -1 used for model building at the start), symmetry transformation (2 - x, 1 - y, 1 - z) for some atom coordinates applied, introducing a split position for Co₂ to get the complete Λ and Δ forms of the inversion related chiral complexes separated, using the information of disorder refinement from the *SHELXL* guide book (Müller *et al.*, 2006).

Figures

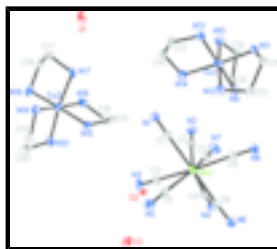


Fig. 1. *ORTEP* view of the title compound showing 30% probability displacement ellipsoids and atom labelling scheme. Hydrogen atoms have been omitted for clarity. The occupancies of the water oxygen atoms are 0.4 (O1), 0.3 (O2), 0.3 (O3).

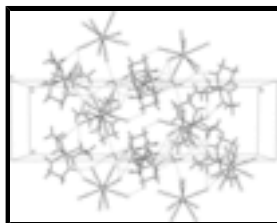
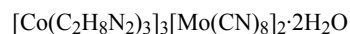


Fig. 2. Packing diagram of the title compound presenting a three-dimensional hydrogen bonding network.

Tris[tris(ethane-1,2-diamine)cobalt(II)] bis[octacyanomolybdate(V)] dihydrate

Crystal data



$M_r = 1361.96$

Monoclinic, $P2_1/n$

Hall symbol: -P 2yn

$a = 9.2113$ (3) Å

$b = 30.5439$ (8) Å

$c = 11.4022$ (3) Å

$\beta = 94.1380$ (10)°

$V = 3199.63$ (16) Å³

$Z = 2$

$F_{000} = 1398$

$D_x = 1.414$ Mg m⁻³

Mo $K\alpha$ radiation

$\lambda = 0.71073$ Å

Cell parameters from 7571 reflections

$\theta = 2.1$ – 26.4 °

$\mu = 1.20$ mm⁻¹

$T = 153$ (2) K

Block, yellow

$0.25 \times 0.23 \times 0.18$ mm

Data collection

Bruker SMART CCD
diffractometer

Radiation source: sealed tube

Monochromator: graphite

6276 independent reflections

4952 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.054$

$T = 153(2)$ K $\theta_{\max} = 26.0^\circ$
 φ and ω scans $\theta_{\min} = 3.0^\circ$
 Absorption correction: multi-scan
 (SADABS; Bruker, 2000) $h = -11 \rightarrow 11$
 $T_{\min} = 0.754$, $T_{\max} = 0.813$ $k = -37 \rightarrow 37$
 27139 measured reflections $l = -14 \rightarrow 14$

Refinement

Refinement on F^2 Secondary atom site location: difference Fourier map
 Least-squares matrix: full Hydrogen site location: inferred from neighbouring sites
 $R[F^2 > 2\sigma(F^2)] = 0.047$ H-atom parameters constrained
 $wR(F^2) = 0.101$ $w = 1/[\sigma^2(F_o^2) + (0.0504P)^2]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $S = 1.06$ $(\Delta/\sigma)_{\max} < 0.001$
 6276 reflections $\Delta\rho_{\max} = 0.42 \text{ e } \text{\AA}^{-3}$
 415 parameters $\Delta\rho_{\min} = -0.55 \text{ e } \text{\AA}^{-3}$
 56 restraints Extinction correction: none
 Primary atom site location: structure-invariant direct methods

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
Mo1	0.40449 (4)	0.336480 (9)	0.62354 (3)	0.03307 (10)	
C1	0.5956 (4)	0.34579 (12)	0.5242 (3)	0.0339 (7)	
N1	0.6910 (3)	0.34961 (9)	0.4732 (3)	0.0329 (6)	
C2	0.3472 (4)	0.37899 (11)	0.4797 (3)	0.0354 (8)	
N2	0.3071 (4)	0.40123 (10)	0.3965 (3)	0.0374 (7)	
C3	0.4699 (4)	0.40112 (12)	0.6844 (3)	0.0380 (8)	
N3	0.5017 (4)	0.43479 (9)	0.7151 (3)	0.0406 (8)	
C4	0.2090 (4)	0.36648 (11)	0.6775 (3)	0.0313 (7)	
N4	0.1008 (4)	0.38188 (9)	0.7010 (3)	0.0385 (7)	
C5	0.5642 (4)	0.33094 (11)	0.7700 (3)	0.0335 (8)	
N5	0.6477 (4)	0.32846 (11)	0.8481 (3)	0.0399 (7)	

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C6	0.3104 (4)	0.29450 (11)	0.7497 (3)	0.0328 (7)	
N6	0.2601 (3)	0.26949 (10)	0.8135 (3)	0.0341 (6)	
C7	0.4994 (4)	0.27234 (11)	0.5986 (3)	0.0327 (7)	
N7	0.5468 (3)	0.23832 (10)	0.5864 (3)	0.0357 (7)	
N8	0.1493 (3)	0.28720 (9)	0.4532 (2)	0.0307 (6)	
C8	0.2437 (4)	0.30333 (11)	0.5112 (3)	0.0327 (7)	
Co1	0.44625 (5)	0.313270 (15)	0.15051 (4)	0.03194 (13)	
N9	0.2527 (3)	0.33083 (9)	0.2003 (3)	0.0327 (6)	
H9C	0.1911	0.3071	0.1969	0.039*	
H9D	0.2616	0.3408	0.2766	0.039*	
C9	0.1923 (4)	0.36582 (11)	0.1216 (3)	0.0349 (8)	
H9A	0.2359	0.3943	0.1464	0.042*	
H9B	0.0854	0.3678	0.1250	0.042*	
C10	0.2286 (4)	0.35532 (12)	0.0022 (3)	0.0323 (7)	
H10A	0.2033	0.3802	-0.0510	0.039*	
H10B	0.1731	0.3293	-0.0271	0.039*	
N10	0.3897 (3)	0.34622 (9)	0.0052 (3)	0.0350 (7)	
H10C	0.4406	0.3721	0.0044	0.042*	
H10D	0.4105	0.3301	-0.0596	0.042*	
N11	0.3701 (3)	0.25818 (9)	0.0810 (3)	0.0345 (7)	
H11C	0.2706	0.2601	0.0678	0.041*	
H11D	0.4087	0.2537	0.0097	0.041*	
C11	0.4077 (4)	0.21994 (11)	0.1607 (3)	0.0327 (7)	
H11A	0.5070	0.2091	0.1497	0.039*	
H11B	0.3375	0.1957	0.1457	0.039*	
C12	0.3977 (4)	0.23912 (11)	0.2871 (3)	0.0324 (7)	
H12A	0.2948	0.2453	0.3011	0.039*	
H12B	0.4361	0.2177	0.3466	0.039*	
N12	0.4821 (3)	0.27916 (9)	0.2972 (2)	0.0304 (6)	
H12C	0.5796	0.2726	0.3091	0.036*	
H12D	0.4554	0.2953	0.3603	0.036*	
N13	0.6390 (3)	0.29952 (9)	0.1000 (2)	0.0300 (6)	
H13C	0.6599	0.2706	0.1156	0.036*	
H13D	0.6395	0.3037	0.0201	0.036*	
C13	0.7505 (4)	0.32704 (10)	0.1603 (3)	0.0333 (8)	
H13A	0.8376	0.3286	0.1145	0.040*	
H13B	0.7797	0.3150	0.2391	0.040*	
C14	0.6848 (4)	0.37129 (11)	0.1710 (3)	0.0351 (8)	
H14A	0.7492	0.3900	0.2230	0.042*	
H14B	0.6716	0.3854	0.0928	0.042*	
N14	0.5431 (3)	0.36574 (9)	0.2211 (3)	0.0345 (7)	
H14C	0.4861	0.3901	0.2053	0.041*	
H14D	0.5563	0.3625	0.3014	0.041*	
Co2	0.9953 (18)	0.5039 (4)	0.4938 (14)	0.0293 (10)	0.50
N15	0.8779 (19)	0.4547 (8)	0.5513 (12)	0.029 (3)	0.50
H15A	0.9260	0.4286	0.5442	0.035*	0.50
H15B	0.8607	0.4588	0.6290	0.035*	0.50
C15	0.7350 (7)	0.4548 (2)	0.4745 (6)	0.0330 (15)	0.50
H15C	0.7487	0.4396	0.3994	0.040*	0.50

H15D	0.6595	0.4389	0.5151	0.040*	0.50
C16	0.6874 (8)	0.5000 (2)	0.4509 (7)	0.0376 (16)	0.50
H16A	0.6584	0.5139	0.5241	0.045*	0.50
H16B	0.6027	0.5003	0.3924	0.045*	0.50
N16	0.8157 (18)	0.5252 (7)	0.4031 (15)	0.037 (3)	0.50
H16C	0.8045	0.5548	0.4136	0.044*	0.50
H16D	0.8208	0.5197	0.3242	0.044*	0.50
N17	1.0281 (19)	0.4680 (8)	0.3505 (16)	0.031 (3)	0.50
H17A	1.0698	0.4407	0.3802	0.037*	0.50
H17B	0.9572	0.4652	0.3117	0.037*	0.50
C17	1.1534 (9)	0.4942 (2)	0.3021 (6)	0.0366 (16)	0.50
H17C	1.2463	0.4872	0.3471	0.044*	0.50
H17D	1.1628	0.4866	0.2186	0.044*	0.50
C18	1.1218 (8)	0.5402 (2)	0.3126 (6)	0.0364 (16)	0.50
H18A	1.0262	0.5469	0.2711	0.044*	0.50
H18B	1.1969	0.5577	0.2759	0.044*	0.50
N18	1.119 (2)	0.5524 (8)	0.4429 (15)	0.046 (6)	0.50
H18C	1.0773	0.5793	0.4531	0.055*	0.50
H18D	1.2108	0.5518	0.4810	0.055*	0.50
N19	1.1687 (17)	0.4768 (7)	0.5739 (13)	0.033 (4)	0.50
H19A	1.1711	0.4470	0.5628	0.040*	0.50
H19B	1.2538	0.4892	0.5523	0.040*	0.50
C19	1.1358 (9)	0.4890 (2)	0.6990 (8)	0.045 (2)	0.50
H19C	1.0650	0.4676	0.7262	0.054*	0.50
H19D	1.2267	0.4858	0.7502	0.054*	0.50
C20	1.0790 (9)	0.5325 (2)	0.7165 (8)	0.047 (2)	0.50
H20A	1.1496	0.5551	0.6956	0.057*	0.50
H20B	1.0564	0.5368	0.7993	0.057*	0.50
N20	0.9413 (19)	0.5345 (9)	0.6346 (16)	0.032 (4)	0.50
H20C	0.9271	0.5615	0.6229	0.038*	0.50
H20D	0.8796	0.5215	0.6781	0.038*	0.50
O1	0.9811 (9)	0.5154 (2)	0.0352 (6)	0.055 (2)	0.40
H1A	1.0629	0.5055	0.0173	0.066*	0.40
H1B	0.9246	0.5179	-0.0268	0.066*	0.40
O2	0.0345 (9)	0.6253 (3)	0.6065 (6)	0.0315 (17)	0.30
H2A	0.0124	0.6518	0.5917	0.038*	0.30
H2B	-0.0393	0.6092	0.5896	0.038*	0.30
O3	0.4984 (11)	0.4328 (3)	0.9766 (8)	0.048 (2)	0.30
H3D	0.4504	0.4525	0.9384	0.058*	0.30
H3C	0.5859	0.4412	0.9908	0.058*	0.30

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Mo1	0.03407 (18)	0.02993 (16)	0.03468 (17)	-0.00118 (12)	-0.00121 (12)	0.00009 (12)
C1	0.0324 (19)	0.0422 (19)	0.0274 (17)	-0.0018 (15)	0.0042 (14)	-0.0041 (15)
N1	0.0323 (16)	0.0348 (15)	0.0314 (15)	-0.0073 (12)	0.0012 (13)	0.0006 (12)
C2	0.042 (2)	0.0281 (16)	0.0366 (19)	-0.0046 (14)	0.0066 (16)	0.0034 (15)

supplementary materials

N2	0.0466 (19)	0.0355 (15)	0.0317 (16)	0.0061 (13)	0.0142 (14)	0.0055 (13)
C3	0.043 (2)	0.0377 (19)	0.0322 (19)	-0.0108 (16)	-0.0009 (16)	-0.0008 (15)
N3	0.0393 (18)	0.0302 (16)	0.0505 (19)	-0.0123 (13)	-0.0100 (14)	-0.0040 (14)
C4	0.037 (2)	0.0283 (16)	0.0288 (17)	0.0002 (14)	0.0031 (14)	0.0001 (14)
N4	0.0437 (19)	0.0296 (14)	0.0425 (18)	0.0004 (13)	0.0057 (14)	-0.0006 (13)
C5	0.0323 (19)	0.0338 (17)	0.0343 (19)	0.0010 (14)	0.0029 (15)	-0.0022 (15)
N5	0.0345 (17)	0.0517 (18)	0.0333 (16)	0.0051 (14)	0.0015 (14)	0.0048 (14)
C6	0.0351 (19)	0.0343 (17)	0.0282 (17)	-0.0028 (15)	-0.0027 (14)	-0.0049 (15)
N6	0.0363 (17)	0.0363 (15)	0.0290 (15)	-0.0033 (13)	-0.0023 (12)	-0.0003 (13)
C7	0.0355 (19)	0.0335 (18)	0.0284 (17)	-0.0013 (14)	-0.0022 (14)	0.0047 (14)
N7	0.0304 (16)	0.0426 (17)	0.0325 (16)	0.0137 (13)	-0.0092 (12)	-0.0091 (13)
N8	0.0246 (14)	0.0328 (14)	0.0336 (15)	-0.0002 (11)	-0.0044 (12)	-0.0045 (12)
C8	0.0338 (19)	0.0347 (17)	0.0297 (17)	0.0036 (14)	0.0041 (14)	-0.0001 (14)
Co1	0.0350 (3)	0.0350 (2)	0.0252 (2)	-0.00011 (19)	-0.00237 (18)	-0.00180 (19)
N9	0.0306 (16)	0.0331 (14)	0.0340 (15)	0.0076 (12)	0.0004 (12)	0.0029 (12)
C9	0.0355 (19)	0.0327 (17)	0.0363 (19)	0.0052 (15)	0.0015 (15)	-0.0089 (15)
C10	0.0300 (18)	0.0378 (18)	0.0282 (18)	0.0065 (14)	-0.0040 (13)	-0.0036 (14)
N10	0.0396 (17)	0.0286 (14)	0.0366 (16)	0.0045 (12)	0.0016 (13)	0.0021 (12)
N11	0.0402 (18)	0.0293 (14)	0.0327 (15)	-0.0047 (12)	-0.0059 (12)	-0.0090 (12)
C11	0.0361 (19)	0.0322 (16)	0.0288 (17)	-0.0001 (14)	-0.0046 (14)	0.0002 (14)
C12	0.0277 (18)	0.0329 (17)	0.0354 (18)	0.0018 (13)	-0.0069 (14)	0.0068 (14)
N12	0.0320 (15)	0.0283 (13)	0.0295 (14)	0.0054 (11)	-0.0066 (11)	0.0006 (12)
N13	0.0299 (15)	0.0325 (14)	0.0272 (14)	0.0038 (11)	-0.0016 (11)	0.0003 (12)
C13	0.0350 (19)	0.0296 (16)	0.0343 (18)	-0.0158 (14)	-0.0041 (14)	0.0134 (14)
C14	0.037 (2)	0.0371 (18)	0.0306 (18)	-0.0040 (15)	-0.0026 (15)	0.0032 (15)
N14	0.0287 (16)	0.0349 (15)	0.0392 (17)	0.0006 (12)	-0.0018 (13)	0.0020 (13)
Co2	0.0417 (17)	0.014 (3)	0.033 (3)	-0.001 (2)	0.0104 (15)	0.007 (2)
N15	0.033 (7)	0.021 (7)	0.034 (6)	-0.008 (4)	0.013 (5)	-0.007 (4)
C15	0.033 (4)	0.027 (3)	0.041 (4)	-0.007 (3)	0.019 (3)	-0.008 (3)
C16	0.033 (4)	0.036 (4)	0.043 (4)	0.001 (3)	-0.005 (3)	-0.003 (3)
N16	0.051 (7)	0.032 (5)	0.025 (6)	0.000 (4)	-0.004 (5)	0.004 (5)
N17	0.015 (7)	0.033 (6)	0.042 (6)	0.002 (5)	-0.011 (5)	-0.014 (5)
C17	0.047 (4)	0.033 (4)	0.031 (4)	-0.001 (3)	0.012 (3)	0.002 (3)
C18	0.027 (4)	0.039 (4)	0.042 (4)	-0.010 (3)	-0.002 (3)	0.011 (3)
N18	0.039 (8)	0.025 (7)	0.078 (11)	-0.012 (4)	0.029 (7)	-0.019 (6)
N19	0.037 (6)	0.035 (6)	0.028 (7)	-0.008 (4)	-0.003 (4)	0.010 (4)
C19	0.040 (4)	0.025 (3)	0.067 (6)	0.002 (3)	-0.014 (4)	0.006 (4)
C20	0.039 (5)	0.042 (4)	0.063 (5)	0.015 (3)	0.014 (4)	0.014 (4)
N20	0.018 (7)	0.033 (4)	0.042 (6)	0.010 (5)	-0.010 (5)	-0.004 (4)
O1	0.058 (5)	0.056 (4)	0.048 (4)	-0.020 (4)	-0.016 (4)	0.013 (3)
O2	0.033 (4)	0.037 (4)	0.024 (4)	-0.004 (3)	0.003 (3)	0.013 (3)
O3	0.055 (6)	0.041 (5)	0.049 (5)	-0.010 (4)	0.011 (4)	0.009 (4)

Geometric parameters (Å, °)

Mo1—C2	2.128 (4)	C14—H14A	0.9900
Mo1—C8	2.142 (4)	C14—H14B	0.9900
Mo1—C4	2.150 (4)	N14—H14C	0.9200
Mo1—C5	2.151 (4)	N14—H14D	0.9200

Mo1—C6	2.155 (4)	Co2—Co2 ⁱ	0.290 (16)
Mo1—C3	2.164 (4)	Co2—N20	1.95 (2)
Mo1—C7	2.172 (3)	Co2—N19	1.966 (12)
Mo1—C1	2.179 (4)	Co2—N18	1.98 (2)
C1—N1	1.095 (5)	Co2—N15	1.990 (18)
C2—N2	1.204 (5)	Co2—N16	1.994 (15)
C3—N3	1.119 (4)	Co2—N17	2.010 (19)
C4—N4	1.151 (5)	N15—C15	1.528 (17)
C5—N5	1.137 (5)	N15—H15A	0.9200
C6—N6	1.173 (5)	N15—H15B	0.9200
C7—N7	1.140 (4)	C15—C16	1.468 (8)
N8—C8	1.163 (4)	C15—H15C	0.9900
Co1—N13	1.952 (3)	C15—H15D	0.9900
Co1—N11	1.968 (3)	C16—N16	1.542 (17)
Co1—N10	1.976 (3)	C16—H16A	0.9900
Co1—N14	1.977 (3)	C16—H16B	0.9900
Co1—N12	1.978 (3)	N16—H16C	0.9200
Co1—N9	1.983 (3)	N16—H16D	0.9200
N9—C9	1.478 (4)	N17—C17	1.540 (18)
N9—H9C	0.9200	N17—H17A	0.968
N9—H9D	0.9200	N17—H17B	0.767
C9—C10	1.460 (5)	C17—C18	1.441 (9)
C9—H9A	0.9900	C17—H17C	0.9900
C9—H9B	0.9900	C17—H17D	0.9900
C10—N10	1.507 (4)	C18—N18	1.534 (16)
C10—H10A	0.9900	C18—H18A	0.9900
C10—H10B	0.9900	C18—H18B	0.9900
N10—H10C	0.9200	N18—H18C	0.9200
N10—H10D	0.9200	N18—H18D	0.9200
N11—C11	1.505 (4)	N19—C19	1.526 (12)
N11—H11C	0.9200	N19—H19A	0.9200
N11—H11D	0.9200	N19—H19B	0.9200
C11—C12	1.565 (5)	C19—C20	1.445 (8)
C11—H11A	0.9900	C19—H19C	0.9900
C11—H11B	0.9900	C19—H19D	0.9900
C12—N12	1.450 (4)	C20—N20	1.522 (11)
C12—H12A	0.9900	C20—H20A	0.9900
C12—H12B	0.9900	C20—H20B	0.9900
N12—H12C	0.9200	N20—H20C	0.845
N12—H12D	0.9200	N20—H20D	0.876
N13—C13	1.460 (4)	O1—O1 ⁱⁱ	1.299 (16)
N13—H13C	0.9200	O1—H1A	0.85
N13—H13D	0.9200	O1—H1B	0.85
C13—C14	1.489 (5)	O2—H2A	0.85
C13—H13A	0.9900	O2—H2B	0.85
C13—H13B	0.9900	O3—H3D	0.85
C14—N14	1.472 (5)	O3—H3C	0.85
C2—Mo1—C8	72.86 (13)	C14—C13—H13A	110.4

supplementary materials

C2—Mo1—C4	78.12 (14)	N13—C13—H13B	110.4
C8—Mo1—C4	79.35 (13)	C14—C13—H13B	110.4
C2—Mo1—C5	139.82 (14)	H13A—C13—H13B	108.6
C8—Mo1—C5	146.29 (13)	N14—C14—C13	107.7 (3)
C4—Mo1—C5	110.56 (13)	N14—C14—H14A	110.2
C2—Mo1—C6	141.99 (14)	C13—C14—H14A	110.2
C8—Mo1—C6	79.84 (13)	N14—C14—H14B	110.2
C4—Mo1—C6	71.15 (13)	C13—C14—H14B	110.2
C5—Mo1—C6	73.77 (13)	H14A—C14—H14B	108.5
C2—Mo1—C3	74.81 (13)	C14—N14—Co1	108.8 (2)
C8—Mo1—C3	141.97 (14)	C14—N14—H14C	109.9
C4—Mo1—C3	74.87 (14)	Co1—N14—H14C	109.9
C5—Mo1—C3	70.36 (14)	C14—N14—H14D	109.9
C6—Mo1—C3	116.44 (13)	Co1—N14—H14D	109.9
C2—Mo1—C7	121.95 (13)	H14C—N14—H14D	108.3
C8—Mo1—C7	76.28 (13)	N20—Co2—N19	93.8 (10)
C4—Mo1—C7	140.82 (13)	N20—Co2—N18	94.0 (9)
C5—Mo1—C7	76.81 (13)	N19—Co2—N18	89.1 (13)
C6—Mo1—C7	74.67 (13)	N20—Co2—N15	85.0 (12)
C3—Mo1—C7	139.42 (14)	N19—Co2—N15	88.3 (8)
C2—Mo1—C1	71.62 (14)	N18—Co2—N15	177.2 (13)
C8—Mo1—C1	107.49 (13)	N20—Co2—N16	91.1 (8)
C4—Mo1—C1	144.87 (13)	N19—Co2—N16	173.6 (15)
C5—Mo1—C1	83.18 (13)	N18—Co2—N16	94.5 (8)
C6—Mo1—C1	143.45 (14)	N15—Co2—N16	88.1 (12)
C3—Mo1—C1	80.26 (14)	N20—Co2—N17	172.8 (15)
C7—Mo1—C1	72.66 (13)	N19—Co2—N17	89.0 (8)
N1—C1—Mo1	178.5 (3)	N18—Co2—N17	92.6 (11)
N2—C2—Mo1	175.7 (3)	N15—Co2—N17	88.4 (8)
N3—C3—Mo1	178.9 (4)	N16—Co2—N17	85.6 (11)
N4—C4—Mo1	176.5 (3)	C15—N15—Co2	106.0 (12)
N5—C5—Mo1	179.1 (3)	C15—N15—H15A	110.5
N6—C6—Mo1	175.8 (3)	Co2—N15—H15A	110.5
N7—C7—Mo1	178.7 (3)	C15—N15—H15B	110.5
N8—C8—Mo1	175.3 (3)	Co2—N15—H15B	110.5
N13—Co1—N11	90.15 (13)	H15A—N15—H15B	108.7
N13—Co1—N10	93.03 (12)	C16—C15—N15	109.8 (10)
N11—Co1—N10	91.88 (12)	C16—C15—H15C	109.7
N13—Co1—N14	84.34 (12)	N15—C15—H15C	109.7
N11—Co1—N14	174.03 (13)	C16—C15—H15D	109.7
N10—Co1—N14	90.71 (13)	N15—C15—H15D	109.7
N13—Co1—N12	92.07 (12)	H15C—C15—H15D	108.2
N11—Co1—N12	85.60 (12)	C15—C16—N16	108.0 (9)
N10—Co1—N12	174.32 (13)	C15—C16—H16A	110.1
N14—Co1—N12	92.28 (12)	N16—C16—H16A	110.1
N13—Co1—N9	176.73 (12)	C15—C16—H16B	110.1
N11—Co1—N9	92.57 (13)	N16—C16—H16B	110.1
N10—Co1—N9	85.07 (12)	H16A—C16—H16B	108.4
N14—Co1—N9	93.01 (12)	C16—N16—Co2	106.4 (12)

N12—Co1—N9	89.96 (12)	C16—N16—H16C	110.4
C9—N9—Co1	109.2 (2)	Co2—N16—H16C	110.4
C9—N9—H9C	109.8	C16—N16—H16D	110.4
Co1—N9—H9C	109.8	Co2—N16—H16D	110.4
C9—N9—H9D	109.8	H16C—N16—H16D	108.6
Co1—N9—H9D	109.8	C17—N17—Co2	99.6 (12)
H9C—N9—H9D	108.3	C17—N17—H17A	106.5
C10—C9—N9	107.8 (3)	Co2—N17—H17A	105.3
C10—C9—H9A	110.1	C17—N17—H17B	118.8
N9—C9—H9A	109.5	Co2—N17—H17B	111.0
C10—C9—H9B	110.2	H17A—N17—H17B	113.9
N9—C9—H9B	110.7	C18—C17—N17	108.5 (10)
H9A—C9—H9B	108.6	C18—C17—H17C	110.0
C9—C10—N10	108.1 (3)	N17—C17—H17C	110.0
C9—C10—H10A	110.1	C18—C17—H17D	110.0
N10—C10—H10A	110.1	N17—C17—H17D	110.0
C9—C10—H10B	110.1	H17C—C17—H17D	108.4
N10—C10—H10B	110.1	C17—C18—N18	109.5 (11)
H10A—C10—H10B	108.4	C17—C18—H18A	109.8
C10—N10—Co1	108.2 (2)	N18—C18—H18A	109.8
C10—N10—H10C	110.1	C17—C18—H18B	109.8
Co1—N10—H10C	110.1	N18—C18—H18B	109.8
C10—N10—H10D	110.1	H18A—C18—H18B	108.2
Co1—N10—H10D	110.1	C18—N18—Co2	98.7 (12)
H10C—N10—H10D	108.4	C18—N18—H18C	112.0
C11—N11—Co1	111.1 (2)	Co2—N18—H18C	112.0
C11—N11—H11C	109.4	C18—N18—H18D	112.0
Co1—N11—H11C	109.4	Co2—N18—H18D	112.0
C11—N11—H11D	109.4	H18C—N18—H18D	109.7
Co1—N11—H11D	109.4	C19—N19—Co2	97.0 (10)
H11C—N11—H11D	108.0	C19—N19—H19A	112.4
N11—C11—C12	103.9 (3)	Co2—N19—H19A	112.4
N11—C11—H11A	111.0	C19—N19—H19B	112.4
C12—C11—H11A	111.0	Co2—N19—H19B	112.4
N11—C11—H11B	111.0	H19A—N19—H19B	109.9
C12—C11—H11B	111.0	C20—C19—N19	116.9 (10)
H11A—C11—H11B	109.0	C20—C19—H19C	108.1
N12—C12—C11	108.8 (3)	N19—C19—H19C	108.1
N12—C12—H12A	109.9	C20—C19—H19D	108.1
C11—C12—H12A	109.9	N19—C19—H19D	108.1
N12—C12—H12B	109.9	H19C—C19—H19D	107.3
C11—C12—H12B	109.9	C19—C20—N20	104.3 (12)
H12A—C12—H12B	108.3	C19—C20—H20A	110.9
C12—N12—Co1	108.7 (2)	N20—C20—H20A	110.9
C12—N12—H12C	109.9	C19—C20—H20B	110.9
Co1—N12—H12C	109.9	N20—C20—H20B	110.9
C12—N12—H12D	109.9	H20A—C20—H20B	108.9
Co1—N12—H12D	109.9	C20—N20—Co2	103.6 (13)
H12C—N12—H12D	108.3	C20—N20—H20C	104.4

supplementary materials

C13—N13—Co1	111.2 (2)	Co2—N20—H20C	112.5
C13—N13—H13C	109.4	C20—N20—H20D	100.3
Co1—N13—H13C	109.4	Co2—N20—H20D	117.6
C13—N13—H13D	109.4	H20C—N20—H20D	115.6
Co1—N13—H13D	109.4	H1A—O1—H1B	109.5
H13C—N13—H13D	108.0	H2A—O2—H2B	109.5
N13—C13—C14	106.6 (3)	H3D—O3—H3C	109.5
N13—C13—H13A	110.4		

Symmetry codes: (i) $-x+2, -y+1, -z+1$; (ii) $-x+2, -y+1, -z$.

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N9—H9C \cdots N7 ⁱⁱⁱ	0.92	2.24	3.064 (4)	148
N11—H11C \cdots N7 ⁱⁱⁱ	0.92	2.09	2.985 (4)	165
N12—H12C \cdots N6 ^{iv}	0.92	2.10	2.955 (4)	154
N13—H13C \cdots N8 ^{iv}	0.92	2.55	3.138 (4)	122
N9—H9D \cdots N2	0.92	2.32	3.117 (4)	145
N14—H14D \cdots N1	0.92	2.28	3.130 (5)	154
N10—H10C \cdots O3 ^v	0.92	1.96	2.856 (10)	164
N10—H10D \cdots N5 ^v	0.92	2.49	3.125 (5)	126
N11—H11D \cdots N6 ^v	0.92	2.58	3.163 (4)	122
N13—H13D \cdots N5 ^v	0.92	2.11	3.012 (4)	167
N16—H16C \cdots N4 ^{vi}	0.92	2.52	3.19 (2)	130
N17—H17A \cdots N2 ^{vii}	0.97	2.49	3.29 (2)	140
N20—H20C \cdots N2 ^{vi}	0.85	2.44	3.02 (2)	127
O2—H2A \cdots N8 ^{viii}	0.85 (14)	2.41 (14)	3.207 (9)	156 (12)
O2—H2B \cdots N2 ^{viii}	0.85 (15)	2.50 (15)	3.245 (9)	147 (12)

Symmetry codes: (iii) $x-1/2, -y+1/2, z-1/2$; (iv) $x+1/2, -y+1/2, z-1/2$; (v) $x, y, z-1$; (vi) $-x+1, -y+1, -z+1$; (vii) $x+1, y, z$; (viii) $-x, -y+1, -z+1$.

Fig. 1

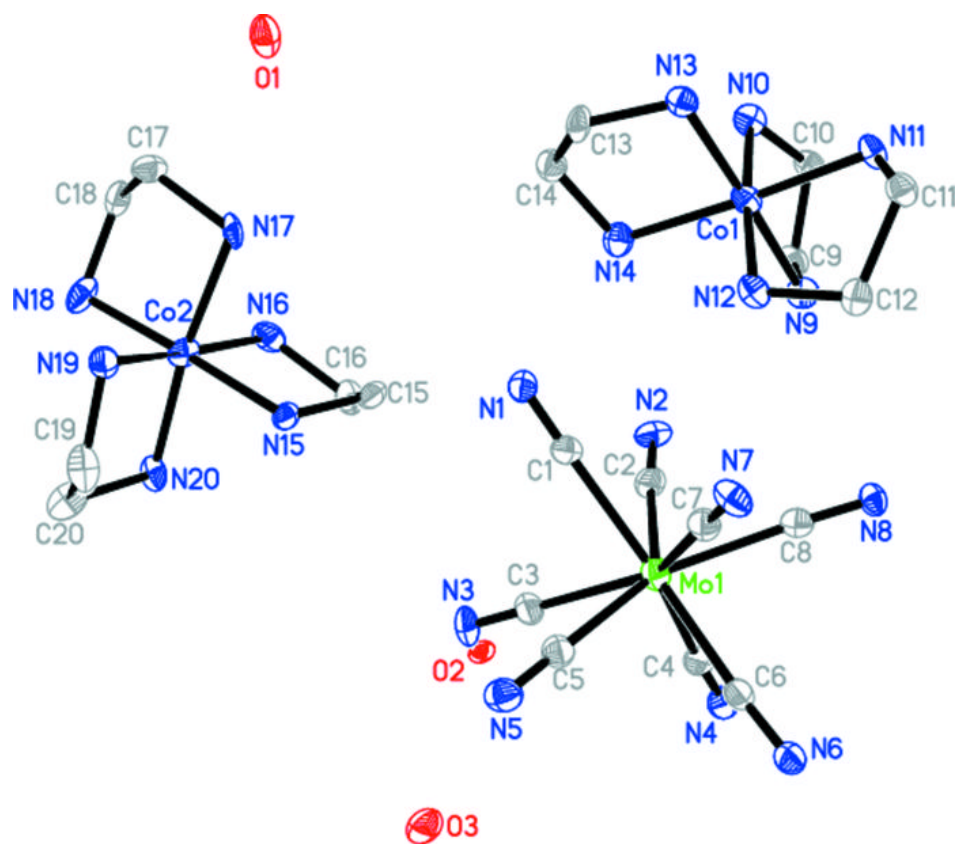


Fig. 2

